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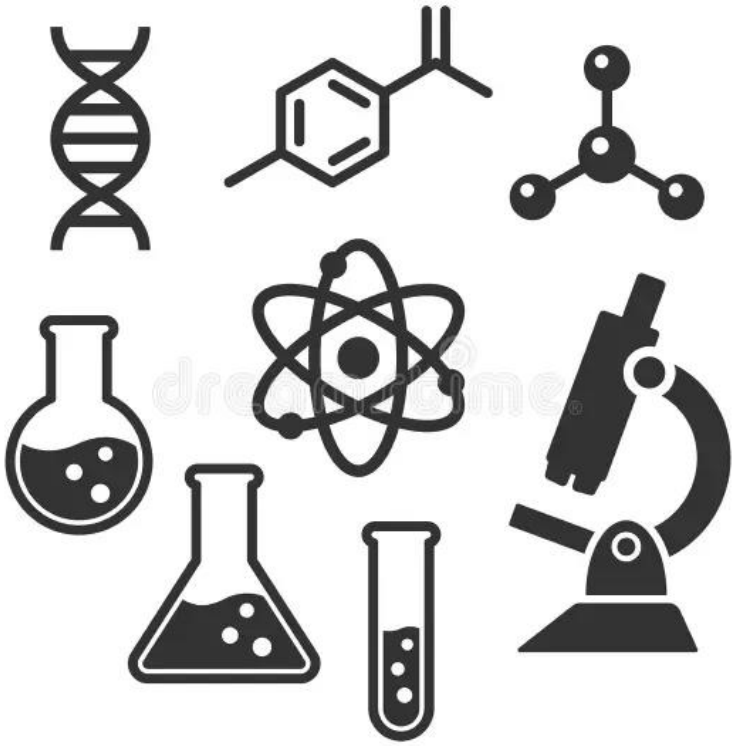
CHEMISTRY CLASS 9TH FBISE MCQS



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CHEMISTRY
CLASS 9TH
FBISE
MCQS

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CHEMISTRY CLASS 9TH (FBISE)
Multiple Choice Questions (MCQs)

1. Example of polyatomic molecule is.....
a) H2SO4
b) S8
c) C6H12O6
d) all of these

2. Identify the element which is in liquid state.
a) Carbon
b) Neon
c) Bromine
d) Magnesium

3. A substance whose atoms have the same atomic number is called.....
a) substance
b) element
c) mixture
d) compound

4. The sum of protons and neutrons in the nucleus of an atom is called.....
a) atomic Number
b) mass number
c) formula Mass
d) atomic Mass

5. Theory of dual nature was presented in.....
a) 1934
b) 1924
c) 1834
d) 1944

6. Molecular formula of Benzene is.....
a) C6H
b) C6H6
c) CHO
d) CH2O

7. 1 gram atom mass of hydrogen = ...
a) 1.008 g
b) 1 mole
c) a & b both
d) none of these

8. Relationship between physical and chemical properties of substances is studied under.....
a) inorganic Chemistry
b) organic Chemistry
c) physical Chemistry
d) biochemistry

9. C6H12O6 is chemical formula of.....
a) starch
b) glucose
c) carbohydrates
d) none of these

10. Natural elements are.....
a) 118
b) 98
c) 84
d) 92

11. When atomic mass is expressed in grams, it is called.....
a) gram atomic mass
b) gram molecular mass
c) gram formula mass
d) molecular mass

12. Hg is symbol of..... element.
a) mercury
b) tungsten
c) silver
d) lead

13. Tungsten is represented by.....
a) T
b) Tn
c) W
d) Sn

14. In water fixed ratio of hydrogen and oxygen by mass is.....

a) 8:1 b) 2:16

c) 1:8 d) 2:1

15. The formula mass of an ionic compound expressed in grams is called.....

a) gram formula mass b) gram formula

c) mole d) All of these

16. Negative charge is carried by.....

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a) free radical b) cation

c) aniond) atom

17. Actual number of atoms are shown in..... formula.

a) Empirical formula b) Molecular formula

c) a & b both d) none of these

18. Triatomic molecules consist of .....atoms.

a) three b) four

c) two

d) more than three

19. When formula mass is expressed in grams is called.....

a) gram formula mass b) gram atomic mass

c) gram molecular mass d) none of these

20. 1 mol of C-12 atoms.....

a) 6.02 × 10-25 b) 6.02 × 10-24

c) 6.02 × 1023 d) 6.02 × 10-23

21. The sum of atomic masses of all the atoms in one formula unit of a substance is called :

a) atomic mass b) mass number

c) formula mass d) atomic mass

22. A specie carrying charge is called:

a) atom b) ion

c) molecule d) all of these

23. Molecular formula = ...

a) n(Empirical formula) b) n/(Empirical formula)

c) (Empirical formula)/n d) none of these

24. 1 amu is equivalent to.....

a) 1.66 × 1023g b) 1.66 × 10-24g

c) 1.66 × 10-23g d) 6.02 × 10-25g

25. Carbon has its atomic number.....

a) 6 b) 8

c) 7 d) 5

26. The common name of Methane is.....

a) coal Gas b) Marsh gas

c) light Gas d) heavy Gas

27. Haemoglobin is an example of .....molecule.

a) mono-atomic b) penta-atomic

c) tetra-atomic d) macro

28. Concerned with the study of synthesis, composition, decomposition and chemical reaction of substances, which take place in the living organisms.

a) Physical Chemistry b) Analytical Chemistry

c) Biochemistry d) Industrial Chemistry

29. Chemical formula has..... types.

a) three b) two

c) one d) four

30. Molecular ion is.....

a) CH4+1 b) CO+1

c) C2H5OH+1 d) all of these

31. The famous book "Al-Asrar" is written by....

a) Jabber Ibn Haiyan b) Ibne-Sina

c) Al Razi d) Al Beruni

32. Chemistry is the study of.....

a) matter b) energy

c) both a and b d) none of these

33. Which of the following chemical process/es was/were introduced by Muslim scientists?

a) Filtration b) fermentation

c) distillation d) all of these

34. Molecular mass of water is.....amu.

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a) 18 b) 12

c) 16 d) 14

35. Which element has atomic number eleven ?

a) Sodium b) Carbon

c) Chlorine d) Oxygen

36. Atom is electrically.....

a) positive b) negative

c) neutral d) none of these

37. Sand is..... mixture.

a) heterogenous b) homogenous

c) soluble d) None of these

38. 1 mole = ...

a) 6.02 × 10-23 units b) 6.02 × 1023 units

c) 6.02 × 1024 units d) 6.02 × 1022 units

39. The element Carbon having the electrons distribution such as .....

a) 1s2,2s1,2p3 b) 1s1,2s2,2p3

c) 1s2,2s2,2p2 d) 1s1,2s1,2p4

40. When an electron jumps from higher energy orbit to lower energy orbit it ..... energy.

a) radiates b) gains

c) a & b both d) no difference occurs

41. Americium-241 is used in.....

a) cancer treatment b) tumors treatment

c) goiter treatment d) smoke detectors

42. "Mg" has an electronic configuration .....

a) 1s2,2s2,2p6,3s1 b) 1s2,2s2,2p6,3s2

c) 1s2,2s2,2p6,3s1 d) 1s2,2s2,2p7,3s2

43. Quantum theory was proposed by.....

a) Rutherford b) Bohr

c) Dalton d) Max Plank

44. d" sub shell carries how many orbitals ?

a) 2 b) 5

c) 3 d) 4

45. The number of neutrons in deuterium are :

a) one b) three

c) two d) four

46. C-14 has neutrons in its nuclei.

a) 6 b) 8

c) 7 d) 9

47. According to Rutherford's model atom, atoms give ..... spectrum.

a) line b) continuous

c) a & b both d) none of these

48. Sodium has electronic distribution such as

a) 1s2,2s2,2p6,3s1 b) 1s1,2s2,2p6,3s2

c) 1s1,2s1,2p6,3s3 d) 1s2,2s1,2p6,3s2

49. If n=2 then it denotes the shell .....

a) L b) M

c) K d) N

50. Rutherford performed an experiment in.....

a) 1913 b) 1914

c) 1911 d) 1917

51. As long as electrons are revolving in fixed circular orbits , they don't .....energy.

a) gain b) lose

c) a & b both d) have

52. John Dalton presented his theory under the title

a) Plum Pudding theory b) Treasure Theory

c) A New system of Chemical theory d) None of these

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53. Rutherford planetary model was similar to.....

a) solar system b) hole system

c) drain system d) none of these

54. Rutherford performed an experiment to determine.....

a) structure of atom b) energy of electron

c) diameter of an atom d) radius of electron

55. Rutherford proposed the model of atom called

a) axial model b) planetary model

c) coaxial model d) all of these

56. Charge of electron is .....

a) 1.6022 × 10-19 C b) 9.87 × 10-19 C

c) 8.11 × 10-31 C d) none of these

57. In Rutherford's experiment alpha-particles were obtained from .....

a) uranium b) polonium

c) radium d) barium

58. Uranium-238 has natural abundance of .....

a) 98.99% b) 78.98%

c) 99.245% d) 87.987%

59. Uranium-235 has natural abundance of .....

a) 65.987% b) 0.56%

c) 87.98% d) 0.75%

60. The energy of the electron in an orbit is proportional to ..... from the nucleus.

a) distance b) closeness

c) a & b both d) none of these

61. Proton is ..... times heavier than an electron.

a) 1856 b) 1865

c) 1837 d) 1975

62. Uranium-238 have ..... neutrons.

a) 143 b) 138

c) 142 d) 146

63. .....is used to measure the moisture content of soil in building construction.

a) Americium-241 b) Iodine-131

c) Carbon -13 d) Californium-252

64. mvr = ...

a) nh/2π b) 2h/nπ

c) nh/2 d) h/2nπ

65. Rutherford concluded that an atom contains ..... in the center of the atom.

a) cork b) space

c) negative portion d) nucleus

66. Orbits are also called .....

a) energy levels b) shells

c) a & b both d) curvatures

67. "Cl" has an electronic configuration.....

a) 1s1,2s3,2p2,3s2,3p3 b) 1s1,2s3,2p4,3s2,3p3

c) 1s1,2s3,2p2,3s2,3p3 d) 1s2,2s2,2p6,3s2,3p5

68. Angular momentum of an electron is an integral multiple of .....

a) hU b) mvr

c) 2h/π d) h/2π

69. Atom is derived from the word .....

a) Atomus b) Atomos

c) Atumous d) none of thses

70. If n=1 then it denotes the shell .....

a) K b) M

c) L d) N

71. "s" sub shell can accommodate a maximum of ..... electrons.

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a) 2 b) 10

c) 6 d) 14

72. Isotopes of uranium have ..... protons.

a) 92 b) 93

c) 143 d) 99

73. The sub shells are represented by.....

a) s, m, k, l b) s, p, d, f

c) K, L, M, N d) all of these

74. Estimation of the age of the plants and animals by radioactive decay is called .....

a) fractional distillation b) carbon dating

c) destructive distillation d) none of these

75. "s" sub shell carries ..... orbitals.

a) 1 b) 3

c) 2 d) 5

76. Rutherford bombarded the gold foil with.....

a) alpha particles b) beta particles

c) gamma rays d) electrons

77. Atom is neutral because .....

a) neutrons = electrons b) electrons = protons

c) neutron = protons d) all of these

78. "p" sub shell can accommodate a maximum of ..... electrons.

a) 14 b) 6

c) 8 d) 10

79. In Rutherford's experiment .....particles were bounced back.

a) 05 b) 02

c) 14 d) 09

80. Bohr considered .....as model.

a) cork b) sun

c) moon d) hydrogen atom

81. Density of water at 0oC is g/cm3

a) 1.000 b) 1.105

c) 2,654 d) 2.876

82. 1st orbit has capacity of electrons .....

a) 8 b) 18

c) 2 d) 32

83. If electrons radiate energy continuously then ..... spectrum will be obtained.

a) continuous b) line

c) a & b both d) none of these.

84. All matter can be ultimately broken down into tiny particles. These particles were named ...

a) compounds b) atoms

c) molecules d) all of these

85. Those elements which show variable valency are called .....

a) representative elements b) transition elements

c) actinides d) lanthanides

86. Elements in a.....have similar chemical properties:

a) period b) group

c) both a and b d) none of these

87. Elements located in the upper right corner of the table will be .....

a) more electronegative b) least electronegative

c) a & b both d) none of these

88. The first alkaline earth metal is .....

a) Be b) Sr

c) Mg d) Ca

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89. The element of the fourth period that is also in group VB is.....

a) Br b) V

c) Pa d) Lu

90. 7th period containing transition elements.....

a) 18 b) 16

c) 10 d) 36

91. Which of the following elements would be similar to carbon:

a) nitrogen b) boron

c) oxygen d) silicon

92. Atomic size of an atom is expressed in terms of .....

a) atomic radii b) covalent radii

c) ionic radii d) all of these

93. 2nd & 3rd periods are also called ..... period

a) long b) very long

c) shortest d) short

94. Noble gases are chemically .....

a) inert b) reactive

c) a& b both d) none of these

95. J.Newlands classified elements in order of their increasing .....

a) atomic mass b) atomic number

c) atomicity d) molecular mass

96. Representative elements are .....

a) s-block b) p-block

c) a & b both d) all of these

97. Oxygen's percentage in atmosphere is .....

a) 20.946% b) 27.876%

c) 23.876% d) 78.084%

98. Properties of elements depends upon.....

a) arrangement of electrons b) atomic number

c) a & b both d) atomic mass

99. Shielding effect is due to:

a) neutron b) proton

c) proton and neutron d) inner shell electron

100. Lithium is present by the..... period.

a) 4th b) 3rd

c) 2nd d) 5

101. The size of an atom is .....

a) fixed b) remains same

c) not fixed d) none of these

102. The second actinide metal is .....

a) Fe b) Sr

c) Cu d) Pa

103. The ionization potential ..... in groups from top to bottom.

a) decreases b) increases

c) remains samed) none of these

104. ..... attempted to classify the elements systematically.

a) Lavoisier b) Mendeleev

c) Newlands d) All of these

105. Lanthanides series ranges from Cerium to.....

a) Barium b) Thorium

c) Lawrencium d) Lutetium

106. In 6th period.....elements are representative.

a) 10 b) 12

c) 8 d) 17

107. Ionization energy is measured in.....

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a) Jmole-1 b) kJmole-1

c) a & b both d) none of these

108. The atomic radius in periodic table ..... in groups from top to bottom.

a) increases b) decreases

c) remains samed) none of these

109. Electronegativity.....in groups.

a) decreases b) increases

c) remains samed) none of these

110. Noble gases contains .....electrons in their valance shell except helium.

a) 3 b) 4

c) 5 d) 8

111. .....calculated the electronegativity values of the elements.

a) Newlands b) Einstein

c) Pauling d) none of these

112. Groups are numbered from I to.....

a) VII b) V

c) VI d) VIII

113. Group III to VIII elements are called ..... elements.

a) s-block b) p-block

c) f- block elements d) d-block elements

114. In 7th period ..... elements are representative.

a) 10 b) 12

c) 8 d) 17

115. Total number of rows in the periodic table are

a) seven b) nine

c) eight d) five

116. Oxygen family contains ..... electrons in their valance shell.

a) 3 b) 4

c) 5 d) 6

117. The rarest naturally-occurring element in the earth's crust is.....

a) Sodium b) Astatine

c) Calcium d) Boron

118. The atomic radius is.....to the number of shells.

a) directly proportional b) inversely proportional

c) not related d) all of these

119. The atomic radius is represented by .....

a) rad b) R

c) r d) all of these

120. ..... arranged elements in the form of table firstly.

a) J.W.Dobereiner b) Mendeleev

c) Lother Meyer d) Lavoisier

121. Group III is also called .....

a) alkali metals b) boron family

c) nitrogen family d) carbon family

122. An element has 8 electrons in its valence shell. It is a member of:

a) alkali family b) halogen family

c) noble family d) carbon family

123. Greater the charge,...will be electronegativity.

a) greater b) lower

c) no effect d) none of these

124. The physical and chemical properties of the elements are the periodic function of their.....

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a) atomic number b) molecular mass

c) formula mass d) atomic mass

125. The minimum energy required to remove valence electron to form positive ion is called

a) ionization energy b) electronegativity

c) electron affinity d) all of these

126. Which of the following would have the smallest ionization energy:

a) K b) P

c) S d) Ca

127. J.W.Dobereiner was a ..... chemist

a) German b) English

c) Russian d) Italian

128. ....are good conductors of heat and electricity.

a) metals b) non metals

c) metalloids d) all of these

129. An element X is in group VI of the periodic table. The ion will be represented by:

a) X+ b) X-

c) X-2 d) X+2

130. A.....bond is formed when two atoms are joined together by sharing of electrons.

a) covalent b) metallic

c) ionic d) none of these

131. Metals tends to form .....

a) cations b) anions

c) a & b both d) all of these

132. A triple bond is denoted by a.....lines.

a) single b) double

c) triple d) none of these

133. Every system in universe tends to lower energy to attain .....

a) stability b) instability

c) a & b both d) none of these

134. The attractive force which holds atom together to form a molecule is called .....

a) parallel force b) friction force

c) chemical bond d) all of these

135. In which year electronic theory of valance was given ?

a) 1916 b) 1919

c) 1913 d) 1923

136. Methane is an example of ..... covalent bond

a) single b) double

c) triple d) none of these

137. Covalent compounds are also called .....

a) ionic compounds b) metallic compounds

c) molecular compounds d) none of these

138. The covalent bond which is formed between atoms having same electronegativity is called

a) polar covalent bond b) non polar covalent bond

c) dative covalent bond d) all of these

139. Dative bond is represented by.....

a) single line b) double line

c) triple line d) arrow

140. In Lewis structure unshared electrons are denoted by .....

a) dot b) cross

c) single line d) a & b both

141. An atom with a charge is called:

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a) an electron b) a molecule

c) a metal d) an ion

142. A double bond is denoted by a.....lines.

a) single b) double

c) triple d) none of these

143. The atom which donates the shared pair of electrons is called .....

a) donor b) acceptor

c) a & b both d) none of these

144. Calcium and chlorine react to form .....

a) calcium chlorinate b) calcium chloride

c) calcium chlorate d) all of these

145. Metallic bond is formed due to.....of electrons.

a) losing b) gaining

c) sharing d) none of these

146. Molecules have shapes of .....

a) linear b) triangular

c) tetrahedral d) all of these

147. The chemical bond which is formed due to complete transfer of electron is called.....

a) metallic bondb) ionic bond

c) dative covalent bond d) covalent bond

148. Matter exists in ..... states.

a) three b) four

c) five d) two

149. Which is not an amorphous solid ?

a) rubber b) NaCl

c) plastics d) glass

150. Density is inversely proportional to .....

a) mass b) area

c) length d) volume

151. .....describes the relationship between V and P at constant temperature.

a) Charles's law b) Boyle's law

c) Avogadro's law d) Newton's law

152. Boyle's law constant Kb is equal to.....

a) PV b) P/V

c) V/T d) VT

153. The temperature at which vapour pressure of liquid becomes equal to atmospheric or external pressure called .....

a) freezing point b) enthalpy

c) boiling point d) sublimation point

154. Greater the temperature ..... will be evaporation.

a) higher b) lower

c) much lower d) b & c both

155. Diamond is hard due to .....

a) covalent bond b) ionic bond

c) coordinate covalent bond d) metallic bond

156. Density's formula is .....

a) V/m b) Vm

c) m/V d) m/V2

157. Diamond is the crystalline allotropic form of

a) carbon b) sulphur

c) tin d) copper

158. We can smell onions even when packed in a sealed bag due to .....

a) effusion b) diffusion

c) expansion d) evaporation

159. According to Charles's law, greater the temperature ..... will be volume

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a) higher b) lower

c) has no effect d) none of these

160. Amorphous means .....

a) soft b) hard

c) shapeless d) adhesive

161. Matter exists in three physical states .....

a) solid, liquid , gas b) liquid, gas, plasma

c) solid, liquid, plasma d) solid, liquid, vapour

162. Carbon exists in ..... forms.

a) crystalline allotropic b) non crystalline form

c) amorphous allotropic d) all of these

163. Allotropes of an element always show different ..... properties

a) physical b) chemical

c) a & b both d) none of these

164. he temperature at which one allotrope changes into another is called .....

a) melting point b) boiling point

c) freezing point d) transition temperature

165. The molecules of gases liquids have ..... kinetic energy.

a) same b) different

c) a & b both d) none of these

166. Atmospheric pressure is equal to .....

a) 76 mm of Hg b) 760 mm of Hg

c) 76 dm of Hg d) 760 cm of Hg

167. ..... describes the relationship between V and T at constant pressure.

a) Charles's law b) Boyle's law

c) Avogadro's law d) Newton's law

168. Which statement is correct about gases ?

a) have indefinite shape b) have indefinite volume

c) a & b both d) none of these

169. Evaporation causes .....

a) heating b) cooling

c) smelling d) ignition

170. Carbon exists in the crystalline forms namely

a) diamond b) graphite

c) bucky balls d) all of these

171. Water has .....intermolecular forces than alcohol.

a) lower b) higher

c) much higher d) none of these

172. Which statement is correct about solids ?

a) have definite shape b) have definite volume

c) a & b both d) none of these

173. Alcohol has ..... rate of evaporation than water.

a) higher b) lower

c) both a & b d) much lower

174. Gases have low density than .....

a) liquids b) solids

c) a & b both d) none of these

175. According to Charles's law volume of the gas is ..... to absolute temperature.

a) directly proportional b) inversely proportional

c) has no effect d) none of these

176. Charles's law is applicable when ..... is constant.

a) volume b) pressure

c) temperature d) all of these

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177. The temperature at which solid starts melting is called .....

a) melting point b) boiling point

c) freezing point d) none of these

178. Evaporation depends on .....

a) surface area b) temperature

c) intermolecular forces d) all of these

179. According to Boyle's law, greater the pressure ..... will be volume

a) higher b) lower

c) has no effect d) none of these

180. The conversion of vapors into liquid state is called .....

a) sublimation b) evaporation

c) condensation d) crystallization

181. Phenol has molecular formula of .....

a) C6H5OH b) CHCl3

c) C5H6 d) C2H3

182. Example of solid aerosol is .....

a) airborne particulate matter b) smoke

c) auto exhaust d) all of these

183. Solution which can dissolve further amount of solute at particular temperature is called ...

a) super saturated solution b) saturated solution

c) unsaturated solution d) none of these

184. Liquid aerosol is .....

a) gas dispersed in liquid b) solid dispersed in gas

c) solid dispersed in liquid d) liquid dispersed in gas

185. Table salt dissolved in water is an example of.....

a) solution b) colloids

c) suspension d) none of these

186. Solution which cannot dissolve further amount of solute at particular temperature is called .....

a) super saturated solution b) saturated solution

c) unsaturated solution d) none of these

187. A solution containing more amount of solute than the amount required to saturate at particular temperature is called .....

a) super saturated solution b) saturated solution

c) unsaturated solution d) none of these

188. Fats, oils and greases are ..... in non polar solvents.

a) partially insoluble b) soluble

c) partially soluble d) insoluble

189. The process of settling of particles at bottom is known as .....

a) Tyndall effect b) suspension

c) sedimentation d) eutrophication

190. 1 molar solution of glucose means ..... of glucose per dm3 of solution.

a) 120g b) 24g

c) 180g d) 40g

191. At higher temperature gases are generally.....

a) insoluble b) less soluble

c) soluble d) none of these

192. Types of solutions are .....

a) three b) five

c) seven d) nine

193. Air is an example of ..... solution.

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a) gaseous b) liquid

c) solid d) all of these

194. Increasing temperature usually.....gas solubility.

a) decreases b) increases

c) doesn't effect d) none of these

195. ..... is an example of suspension.

a) table salt b) milk in water

c) sugar in water d) flour in water

196. In.....composition is not uniform throughout

a) suspension b) solution

c) colloid d) none of these

197. Molarity is represented by .....

a) ml b) m

c) M d) all of

198. solution which is made up of two components is called .....

a) saturated solution b) super saturated solution

c) binary solution d) all of these

199. Brass (alloy) is an example of .....solution.

a) solid in liquid b) liquid in gas

c) liquid in liquid d) solid in solid

200. Example of solid in liquid solution is .....

a) sea water b) smoke

c) amalgam d) brass

201. Brass is an alloy of ..... and zinc.

a) copper b) aluminium

c) silverd) gold

202. ..... is heterogeneous.

a) solution b) colloid

c) suspension d) both b & c

203. 10% m/m solution of sugar means .....

a) 10g sugar in 110g solvent b) 100g sugar in 10g solvent

c) 10g sugar in 90g solvent d) 90g sugar in 10g solvent

204. A solution which contains small amount of solute is called .....

a) concentrated solution b) saturated solution

c) dilute solution d) none of these

205. Solid dispersed in liquid is phase of which class of colloid ?

a) gel b) foam

c) sol d) solid aerosol

206. Types of solid solution are .....

a) two b) four

c) three d) five

207. Petrol is ..... in water.

a) soluble b) insoluble

c) partially soluble d) none of these

208. Anions move towards.....in electrolytic cell.

a) anode b) cathode

c) a & b both d) none of these

209. In NaH, the oxidation number of hydrogen is

a) -1 b) 0

c) +2 d) -2

210. The device in which chemical energy is converted into electrical energy is called .....

a) electrolytic cell b) galvanic cell

c) solar cell d) all of these

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211. Sodium metal was first discovered by .....

a) Sir Humphrey Davey b) Berzilius

c) Boyle d) Volta

212. The electrode at which cation gathers and leaves the electron is called .....

a) anode b) cathode

c) a & b both d) none of these

213. Oxidation number of S is .....

a) +5 b) +6

c) 0 d) +2

214. The down cell consists of ..... container.

a) steel b) iron

c) copper d) chromium

215. Manganese have ..... oxidation states.

a) 4 b) 2

c) 5 d) 6

216. In dry cell container is filled with mixture which is in .....form.

a) granular b) paste

c) solution d) none of these

217. The necessary condition for rusting is .....

a) moisture b) fire

c) air d) a & c both

218. The solutions in different containers in Daniel cell are connected through a bridge called ...

a) salt bridge b) acid bridge

c) base bridge d) iron bridge

219. The net chemical changes obtained by adding the two half cell reactions are called .....

a) cell reaction b) half reaction

c) tube reaction d) all of these

220. Electroplating of Tin is carried out in .....

a) electrolytic cell b) galvanic cell

c) a & b both d) Daniel cell

221. bH2, metals, SO2, and CO are .....

a) oxidizing agent b) reducing agent

c) a & b both d) none of these

222. Cations move towards.....in electrolytic cell.

a) anode b) cathode

c) a & b both d) none of these

223. Corrosion is a ..... process.

a) non-spontaneous b) spontaneous

c) a & b both d) none of these

224. The process which involves gain and loss of electrons simultaneously called .....

a) oxidation reaction b) reduction reaction

c) redox reaction d) none of these

225. Chromium solution having +3 oxidation state has .....colour

a) blue b) red

c) greend) orange

226. The anode in down cell is made of .....

a) graphite b) iron

c) chromium d) copper

227. Chromium solution having +6 oxidation state has ..... color

a) blue b) red

c) greend) orange

228. Cathode carries.....charge.

a) positive b) negative

c) no charge d) all of these

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229. The electrode at which anion gathers and leaves the electron is called .....

a) anode b) cathode

c) a & b both d) none of these

230. An electrolytic process in which a thin layer of one metal is deposited on another metal surface is called .....

a) electroplatingb) surfacing

c) pasting d) all of these

231. Electrochemical cells are used to power .....

a) radios b) automobiles

c) flashlights d) all of these

232. Physical method used for recognition of equilibrium is.....

a) Refractrometry b) Polarimetry

c) Spectrophotometry d) All of these

233. The reaction will more either in the forward direction or in the reverse direction to attain

a) speed up b) disturbance

c) equilibrium d) None of thes

234. How many methods are used to recognize the state of equilibrium.

a) One b) Two

c) Three d) Four

235. The reaction in which reactant react form product is.....

a) Forward reaction b) Reverse reaction

c) Irreversible reaction d) All of these

236. Kc has no unit for a reaction in which number of moles of reactant and products are.....

a) equalb) opposite

c) less d) greater

237. When proposed the Law of Mass Action?

a) 1763 b) 1802

c) 1853 d) 1864

238. .....reactions takes place from right to left.

a) Neutralization b) Reverse

c) Forward d) All of these

239. At initial stage, the rate of..... reaction is very fast.

a) reverse b) forward

c) both a & b d) None of these

240. In Haber process, nitrogen reacts with hydrogen and produce .....

a) Nitric acid b) Ammonia

c) Sulphuric acid d) Oxygen

241. Writing the.....of reaction the product is written as numerator and reactants as denominator.

a) equilibrium constant b) products

c) reactants d) catalyst

242. When the ratio of [products]/[Reactants] is greater than Kc and system is not at equilibrium then more.....is required to such equilibrium.

a) reactants b) product

c) heat d) None of these

243. When the value of Kc is very small it shows that.....

a) reaction will go in forward direction b) reaction will go in reverse direction

c) reaction is at equilibrium d) equilibrium will never established

244. Active masses means.....

a) the total mass of reactants b) the total mass of products

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c) the total mass of reactants d) concentration of reactant and products in moles per dm3 in dilute solution.

245. Rf = kf [A] [B] kfis the rate of constant for the.....reaction.

a) reverse b) forward

c) reactant d) product

246. When we know the balanced chemical equation of a...reaction we can write the Kc.

a) irreversible b) neutralization

c) reversible d) All of these

247. In a chemical reaction the substances that are combine are.....

a) reactants b) product

c) mixture d) atom

248. Rr = Kr [C] [D] Kr is the rate constant for the.....reaction.

a) reverse b) forward

c) concentration d) All of these

249. According to the law of mass action, rate of forward reaction is..... proportional to the product of the concentration of the reactants.

a) inversely b) directly

c) Both a & b d) None of these

250. In a chemical reaction the substances that are formed are.....

a) reactants b) atom

c) products d) None of these

251. The value of equilibrium constant can be calculated of we know the.....of reactants and produce in equilibrium mixture.

a) concentration b) suspension

c) moles d) compound

252. If the condition of the.....are charged, a change in the system may be observed.

a) reactant b) product

c) both a & b d) None of these

253. According to Arrhenius..... is a chemical substance that dissociate in aqueous solution to give hydroxide ion.

a) acid b) base

c) both a & b d) None of thes

254. he acid which produce one proton per molecule are called.....

a) mono-acid base b) poly-acid base

c) mono-protic acid d) poly-protic acid

255. The chemical formula of formic acid is...

a) HCOH b) HCOOH

c) COH d) HOC

256. Sodium is a.....

a) metal b) non-metal

c) compound d) mixture

257. .....are classified in term of the OH- that can be given by per molecule of base when added to water.

a) Acid b) Bases

c) both a & b d) none of these

258. A third classification of acids and bases was introduced in.....

a) 1992 b) 1923

c) 1924 d) 1925

259. .....is also known as washing soda

a) sodium carbonate b) sodium carbonate

c) sodium sulphate d) calcium sulphate

260. A.....consists of positive ion from base and negative ion from an acid.

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a) salt b) lime stone

c) paints d) mirror

261. Solution of soluble salt is mixed and a result...

a) insoluble salt b) soluble salt

c) both a and b d) none of these

262. Acid react with base and neutralized the effect of each other to produce.....and water.

a) acid b) salt

c) base d) both a and b

263. An.....can tell us whether a substance is an acid, base or neutral

a) reaction b) product

c) salt d) indicator

264. The reaction in which two water molecules produce ions is called .....

a) self-ionization b) auto-ionization

c) both a and b d) none of these

265. The table salt is.....

a) NaCl b) Na2SO4

c) Na2CO3 d) KNO3

266. .....is the method of preparation of soluble salt.

a) neutralization b) addition

c) subtraction d) combustion

267. The word acid is derived from a...word.

a) Greek b) Latin

c) Both a & b d) None of these

268. .....is a laboratory method of quantitative chemical analysis.

a) titration b) sublimation

c) condensation d) evaporation

269. A neutral solution at 25C0 has a [O-H] of

a) 1.0 ´ 10-7 M b) 1.0 ´ 107 M

c) 0.1 ´ 10-7 M d) 0.1 ´ 107 M

270. There are.....types of base

a) two b) three

c) four d) five

271. The pH value of detergent solution is approximately.....

a) 9 b) 10

c) 11 d) 12

272. CuO is a.....

a) oxide of nitrogen b) oxide of carbon

c) oxide of oxygen d) metal oxide

273. .....are group of atoms that are formed when one hydrogen atom is removed from an alkane molecule.

a) alkanes b) alkenes

c) alkynes d) alkyl radicals

274. Molecular formula of pentyl is.....

a) C4H9 b) C5H11

c) C3H7 d) C2H5

275. The clothes we wear are made up of.....

a) wood b) plast

c) fibers d) iron

276. Natural gas is used as raw material in the production of.....

a) plastic b) detergents

c) fertilizer d) medicines

277. Carbohydrates is present in.....

a) apple b) beans

c) potatoes d) both a and c

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278. Organic compounds that contain only hydrogen and carbon are called.....

a) carbons b) hydrogens

c) hydrocarbons d) oxides of carbons

279. The molecular formula of decane is.....

a) C10H22 b) C4H22

c) C8H22 d) C10H20

280. Alkanes are.....reactive due to their saturated nature.

a) more b) less

c) non d) all of these

281. Naturally occurring organic compounds are

a) coal b) food

c) Both a & b d) None of these

282. Ali-cyclic organic compounds differ from open chain organic compound by.....

a) O2 b) C2

c) N2 d) none of these

283. The.....is found in combination with or without petroleum reservoirs.

a) coal b) mineral

c) zinc d) natural gas

284. Coal, petroleum and natural gas are collectively called.....

a) rock b) fossil fuel

c) peat d) lignite

285. Alkyl radical is represented generally by a letter.....

a) P b) Q

c) R d) H

286. The important member of aromatic organic compound is.....

a) propene b) butene

c) benzene d) pentene

287. Organic compounds have generally low.....

a) melting point b) boiling point

c) both a & b d) none of these

288. Organic compounds are mainly.....in nature with few exception.

a) polar b) non-polar

c) less polar d) more polar

289. Carbon atom have the ability to covalently bond with other carbon atoms. This self linking ability of carbon is called.....

a) catenation b) aromatization

c) both b & c d) none of these

290. Plants contains organic compounds are.....

a) carbohydrates b) proteins

c) fats d) all of these

291. Natural gas is used as a fuel in automobiles as

a) C.OG b) CGN

c) GNC d) CNG

292. Carbon can make bond with many other types of atoms such as.....

a) hydrogen b) oxygen

c) nitrogen d) all of these

293. The formula of potassium hydroxide is .....

a) OH b) KOH

c) HBr d) HI

294. The simplest alkyne is .....

a) ethane b) ethene

c) ethyne d) propane

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295. .....are also used in the preparation of detergents and surfactants.

a) hydrocarbons b) noble gases

c) metals d) non-metals

296. The.....are gradually increase with an increase in molecular mass.

a) melting point b) density

c) boiling point d) all of these

297. Plant-based oil reduce the level of ..... cholesterol in our bodies.

a) bad b) good

c) both a & b d) none of these

298. The chemical formula of potassium permanganate is .....

a) KMO4 b) KMnO4

c) MnO4 d) KMnO2

299. A number of useful chemicals obtained from hydrocarbons are .....

a) methyl alcohol b) acetic acid

c) choroform d) all of these

300. Example of alkyne is .....

a) proane b) popene

c) propyne d) butane

301. .....can be prepared by addition reaction of alkenes and alkynes

a) Alkane b) Cycloalkane

c) both a & b d) none of these

302. ..... is prepared from dehydro halogenation of ethyl chloride.

a) ethane b) ethene

c) ethyne d) ethanol

303. When H2 is added to ethane, the double bond is converted into single bond and produces....

a) ethane b) ethene

c) ethyne d) ethanol

304. Each different compound should have...name.

a) same b) different

c) both a & b d) none of these

305. In a .....reaction, atoms are removed from a molecule to form a double or triple bond.

a) elimination b) combustion

c) addition d) hydrogenation

306. The general formula of alkynes is .....

a) CnH2n+2 b) CnH2n-2

c) CnH2n+1 d) CnH2n

307. The melting and boiling points of alkane increase regularly in the series with.....in molecular size.

a) decrease b) increase

c) remain same d) change continuously

308. Hydrocarbons can be converted into synthetic rubbers as .....

a) ethane b) ethene

c) ethyne d) propene

309. The combustion of ethyne when it is mixed with pure oxygen produces large amount of heat used for .....

a) welding b) boiling

c) freezing d) all of these

310. An alkane in the shape of the ring is called....

a) alkane b) alkene

c) cycloalkane d) alkene

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311. Which of the following is organic solvents...

a) benzene b) acetone

c) ather d) all of these

312. Alkane that do not contain ring have general formula .....

a) CnH2n+1 b) CH2n+1

c) CnH2n+2 d) CH2n+2

313. The addition of hydrogen to alkenes readily occurs in the presence of ..... as a catalyst.

a) Na b) Br

c) I d) Ni

314. In hydrogenation reaction, the unsaturated hydrocarbon are converted into ..... hydrocarbons.

a) unsaturated b) saturated

c) both a & b d) none of these

315. The general formula of alkene is .....

a) CnH2n+1 b) CnH2n+2

c) CnH2n d) CnH2n-1

316. .....are soluble in non-polar solvents.

a) alkane b) alkene

c) alkyne d) all of these

317. Animal-based fats contain more saturated.....

a) protein b) carbohydrate

c) fats d) all of these

318. Monosaccharide molecules can rotate the plane polarized light to.....

a) left b) right

c) up d) down

319. Fats and oil are used in the manufacture of...

a) soap b) cosmetics

c) paints d) all of the

320. Proteins like.....are chemical regulators in the body.

a) hormons b) enzymes

c) both a & b d) none of thes

321. DNA structure was discovered by.....

a) J.Watson b) F.Crick

c) both a & b d) none of these

322. Cellulose is obtained from.....

a) plants b) animals

c) woodd) all of these

323. Vitamin D regulates blood.....necessary for proper bones and tooth growth.

a) magnesium b) oxygen

c) calcium d) lithium

324. .....produces glucose and glucose.

a) maltose b) lactose

c) sucrosse d) glucose

325. Disulphide bridges are found only in..... containing protein molecules.

a) oxygen b) sulphur

c) nitrogen d) carbon

326. The general formula of sucrose is.....

a) C12H22O10 b) C12H22O11

c) C10H22O10 d) C6H12O6

327. Vitamin C is necessary for.....

a) healing wounds b) bones

c) teeth d) antioxidant

328. All are present in DNA except.....

a) ribose sugar b) phosphate unit

c) nitrogen base d) deoxyribose sugar

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329. The long chain fatty acid of esters are made up of.....fatty acid.

a) 3 b) 4

c) 2 d) 5

330. Majority of changes in DNA are repaired by special.....in the cell.

a) vitamin b) both a & b

c) enzymes d) none of these

331. Polysaccharides are.....

a) sweet b) tasteless

c) bitter d) sour

332. Nucleic acid was discovered by.....

a) Arrhenius b) Guldberg

c) Johannes Friedrich Micscher d) Lewis

333. .....is pentahydoxy ketone.

a) glucose b) fructose

c) chitin d) both a & b

334. Most proteins obtained from animals sources contain all the necessary.....in enough amounts.

a) amino acids b) oxygen

c) nitrogen d) carbon

335. Changes in DNA molecules may be caused by mutagens like.....

a) chemical agents b) radiation

c) viruses d) all of these

336. Deficiency of vitamin A cause.....

a) night blindness b) dry skin

c) both a & b d) none of these

337. Vitamin K is present in.....

a) fresh green vegetables b) bread

c) butler d) corn

338. Protein molecules may be.....

a) fibrous b) globular

c) both a & b d) none of these

339. If all ozone around the earth is condensed, it will form a layer of .....thickness on earth surface.

a) 1.3mm b) 1.4mm

c) 2.4 mm d) 2.5 mm

340. At pH....., most fish egg cannot hatch.

a) 4 b) 5

c) 6 d) 7

341. The calcium sulphate is.....in H_2O

a) soluble b) insoluble

c) both a & b d) none of these

342. The pollutants can be in the form of .....

a) solid particles b) liquid particles

c) gases d) all of these

343. The first layer of atmosphere, which is closed to the surface of earth is called.....

a) thermosphere b) mesosphere

c) stratosphere d) troposphere

344. The molecular formula of ozone is.....

a) O_2 b) N_2

c) SO_2 d) O_3

345. The auto exhaust emission is directly contributing to the.....pollution.

a) air b) noise

c) water d) all of these

346. The decrease in the number of plants increases the amount of CO_2 in air.

a) decrease b) increase

c) both a & b d) none of these

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347. Higher concentration of.....causes leaf damage and affects the photosynthesis activities of plants.

a) SOx b) COx

c) NOx d) CO

348. The average height of troposphere from earths surface is about.....

a) 11km b) 12km

c) 13 km d) 14 km

349. CO combines with haemoglobin and reduces the .....carrying capacity of blood.

a) carbon b) oxygen

c) hydrogen d) helium

350. The contribution of natural source towards the particulate emission is ..... than man-made sources.

a) equal b) greater

c) less d) none of these

351. Atmosphere contain.....which is necessary for all living things.

a) oxygen b) calcium

c) magnesium d) sodium

352. An important example of secondary pollutant is.....

a) volatile organic compounds b) ground level ozone

c) volcanic eruption d) both a & b

353. The atomic oxygen is very important and consumes about ..... of ozone found in atmosphere.

a) 16% b) 17%

c) 18% d) 19%

354. VOCs such as methane etc. are produced wherever there are.....

a) fossil fuels burning b) mining

c) extraction d) all of these

355. VOCs contribute.....of natural methane in the environment.

a) 75% b) 76%

c) 77% d) 78%

356. The depletion of ozone layer allows the entering of UV from the sun into the earth's

a) atmosphere b) stratosphere

c) mesosphere d) thermosphere

357. Frog are sensitive and may not survive below pH .....

a) 3.6 b) 4.6

c) 5.6 d) 6.6

358. .....allows the visible & UV rays from the sun to enter into atmosphere.

a) CO2, NO2 b) O3

c) CH4 d) all of these

359. .....are gases and particles in the atmosphere that harm organisms and effect climate.

a) germs b) air pollutants

c) virus d) bacteria

360. The main source of emission of nitrogen are..

a) bacteria b) microorganisms

c) both a & b d) none of these

361. Atmosphere helps in.....

a) flow of energy b) radio communication

c) movement of aircrafts d) all of these

362. Gaseous pollutants are .....

a) oxides of sulphur b) VOCs

c) oxides of nitrogen d) all of these

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363. The rain having pH less than .....is called acid rain.

a) 5.6 b) 5.7

c) 5.8 d) 5.9

364. Leachate finds its way to rivers and streams through.....

a) rain water b) ozone depletion

c) eutrophication d) green house effects

365. Water reacts with SO3 forming.....

a) H2SO4 b) HNO2

c) HNO3 d) CaCO3

366. ..... proved experimentally that water is not an element but is a compound.

a) Lewis b) Bronsted

c) Henry Cavendish d) Guldberg

367. ...is one of the most common waterborne diseases.

a) blood pressure b) dysentery

c) headache d) TB

368. Water react with NO2 to produce .....

a) HNO2 b) HNO3

c) both a & b d) H2SO4

369. The chemical formula of carbonic acid is.....

a) HNO3 b) CaCO3

c) H2CO3 d) H2SO4

370. The industrial discharge are absorbed by lover life and passes into food chain causing

a) death b) birth defects

c) mental problems d) all of these

371. The chemical formula of sodium stearate is.....

a) Ca++ b) C17H35COO

c) C17H35COONa d) (C17H35COO)2Ca

372. Water react with .....forming carbonic acid.

a) NO2 b) SO2

c) HCl d) CO2

373. Aresenicosis is a disease caused by drinking water with high level of .....

a) fluoride b) arsenic

c) bromine d) iodine

374. Agriculture chemicals which pollute the water are .....

a) insecticides b) pesticides

c) fertilizers d) all of these

375. The latent heat of fusion of ..... is 6kj/mole.

a) CO2 b) NO2

c) SO2 d) H2O

376. Earth has been called the ..... planet.

a) pink b) yellow

c) black d) blue

377. Parasitic infection cause.....

a) ulcer b) fever

c) vomiting d) all of these

378. The purposes of extracting a substance from a solid material that is dissolved in a liquid is called .....

a) chromatography b) bleaching

c) leaching d) all of these

379. The general formula of sodium hydroxide is...

a) NaOH b) KOH

c) CaOH d) CaO

380. Water is a universal.....

a) solute b) solvent

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c) solution d) suspension

381. Some.....use water for cooling purposes.

a) industries b) power station

c) peoples d) both a & b

382. The ..... ions are present in water.

a) Ca+2b) Mg+2

c) Na+2d) both a & b

383. The conductivity of water.....as electrolyte (salt) dissolved in it.

a) decrease b) increases

c) maintain d) none of these

384. Cholera symptoms are usually .....

a) watery diarrhea b) vomiting

c) both a & b d) none of these

385. In children, mercury causes.....

a) brain damage b) learning defects

c) incomplete mental development d) all of these

386. Water has a high dielectric constant of 80 at

a) 160Cb) 170C

c) 180Cd) 200C

387. Stem react with red-hot iron to form the magnetic.....

a) nitrogen oxide b) iron oxide

c) carbon dioxide d) zinc oxide

388. Solid waste is also called.....

a) urban waste b) domestic waste

c) both a & b d) wastewater

389. .....is domestically well known as a water softener.

a) NaCaO b) NH4OH

c) Na2CO3 d) NaOH

390. Most organic fertilizer are derived from plant and animals like.....

a) manure b) bone

c) blood d) all of these

391. When .....is pumped out from the grounds it is mixture of a large number of compounds.

a) paints b) wood

c) waterd) crude oil

392. The petroleum is mainly composed of .....

a) woodb) hydrocarbons

c) soil d) stone

393. Solvay process os used for the commercial preparation of.....

a) NH3 b) Na2CO3

c) (NH2)2CO d) NaOH

394. Magnetic separation method is used for the separation of .....from non-magnetic ore by using magnetic separator.

a) magnetic ore b) non-magnetic ore

c) both a & b d) none of these

395. When the .....pocked is pierced the gas pressure, forces the oil out.

a) juice b) paints

c) iron d) oil

396. The inorganic origin theory of petroleum was put forward by .....chemist.

a) Russian b) Pakistani

c) American d) Japanese

397. Synthetic fertilizers are easy to apply in .....from.

a) granular b) liquid

c) gas d) both a & b

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398. Electrorefining of metal is a process in which obtaining of pure metal from ore by a process of .....

a) electrolysis b) roasting

c) conduction d) filtration

399. The ores are concentrated to sufficient degree of purify by simply picking it with hand and breaking the rock stone with hammer is called

a) hand picking b) froth floatation

c) hydraulic washing d) crushing

400. The general formula of urea is .....

a) NH2CONH2 b) NCONH2

c) NH3CO d) NH4Cl

401. Urea serves as an important role in the metabolism of .....containing compounds by animals.

a) N b) S

c) Cl d) Br

402. Ammonium carbonate reacts with NaCl and form.....

a) NaHCO3 b) NH4Cl

c) both a & b d) Na2CO3

403. The process in which the oxide ore in the fused is reduced with reducing agents such as coke to get metal is called.....

a) roasting b) smelting

c) extraction d) distillation

404. The general formula of iron silicate is .....

a) FeO b) FeS

c) FeSiO3 d) CuO

405. The cuprous sulphide in the blast furnace also oxidizes and form.....

a) Cu2O b) FeS

c) FeO d) SO2

406. For fractional distillation petroleum is heated above .....in a furnace.

a) 4000C b) 2000C

c) 3000C d) 6000C

407. ..... Slag.

a) iron silicate b) Silica

c) calcium carbonate d) Iron Oxide

408. The metal present in combined state is called

a) mineral b) metal

c) solid d) alloy

409. In diesel oil, the number of carbon atoms per molecule is .....

a) C12 to C14 b) C12 to C16

c) C12 to C20 d) C14 to C25

410. CO2 can be obtained from.....

a) NH3 b) CH4

c) CaO d) NaHCO3

411. .....synthesized urea.

a) Ernest Solvay b) Friedrich Wohler

c) Bronsted d) none of these

412. Dry sodium bicarbonate is heated in rotator furnace called.....

a) brain solution b) ammonical solution

c) slaked lime d) calciner

413. Urea is one of the most important...fertilizers.

a) carbon b) nitrogenous

c) sulphur d) halogens

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ANSWER KEY

1. D	2. C	3. B	4. B	5. B	6. B	7. C	8. C	9. B	10. D
11. A	12. A	13. C	14. C	15. A	16. C	17. B	18. A	19. A	20. C
21. C	22. B	23. A	24. B	25. A	26. B	27. D	28. C	29. A	30. D
31. C	32. C	33. D	34. A	35. A	36. C	37. A	38. B	39. C	40. A
41. D	42. B	43. D	44. B	45. A	46. B	47. B	48. A	49. A	50. C
51. B	52. C	53. A	54. A	55. B	56. A	57. C	58. C	59. D	60. A
61. C	62. D	63. A	64. A	65. D	66. C	67. D	68. D	69. B	70. A
71. A	72. A	73. B	74. B	75. A	76. A	77. B	78. B	79. B	80. D
81. A	82. C	83. A	84. B	85. B	86. B	87. A	88. A	89. B	90. C
91. D	92. D	93. D	94. A	95. A	96. C	97. A	98. C	99. D	100. C
101. C	102. D	103. A	104. B	105. D	106. C	107. B	108. A	109. A	110. D
111. C	112. D	113. D	114. C	115. A	116. D	117. B	118. A	119. C	120. B
121. B	122. C	123. A	124. A	125. A	126. A	127. A	128. A	129. C	130. A
131. A	132. C	133. A	134. C	135. A	136. A	137. C	138. B	139. D	140. A
141. D	142. B	143. A	144. B	145. C	146. D	147. B	148. B	149. B	150. D
151. B	152. A	153. C	154. A	155. A	156. C	157. A	158. B	159. A	160. C
161. A	162. D	163. A	164. D	165. B	166. B	167. A	168. C	169. B	170. D
171. B	172. C	173. A	174. C	175. A	176. B	177. A	178. D	179. B	180. C
181. A	182. D	183. C	184. D	185. A	186. B	187. A	188. B	189. C	190. C
191. B	192. D	193. A	194. A	195. D	196. A	197. C	198. C	199. D	200. C
201. A	202. D	203. C	204. C	205. C	206. A	207. B	208. A	209. A	210. B
211. A	212. B	213. C	214. A	215. D	216. B	217. D	218. A	219. A	220. A
221. B	222. B	223. B	224. C	225. C	226. A	227. D	228. B	229. A	230. A
231. D	232. D	233. C	234. B	235. A	236. A	237. D	238. B	239. B	240. B
241. A	242. A	243. B	244. D	245. B	246. C	247. A	248. A	249. B	250. C
251. A	252. C	253. B	254. C	255. B	256. A	257. B	258. B	259. A	260. A
261. A	262. B	263. D	264. C	265. A	266. A	267. B	268. A	269. A	270. B
271. B	272. D	273. D	274. B	275. C	276. C	277. D	278. C	279. A	280. B
281. C	282. D	283. D	284. B	285. C	286. C	287. C	288. B	289. A	290. D
291. D	292. D	293. B	294. C	295. A	296. D	297. A	298. B	299. D	300. C
301. A	302. B	303. A	304. B	305. A	306. B	307. B	308. C	309. A	310. C
311. D	312. C	313. D	314. B	315. C	316. D	317. C	318. B	319. D	320. C
321. C	322. A	323. C	324. A	325. B	326. B	327. A	328. A	329. A	330. C
331. B	332. C	333. B	334. A	335. D	336. C	337. A	338. C	339. D	340. B
341. B	342. D	343. D	344. D	345. A	346. A	347. A	348. B	349. B	350. B
351. A	352. B	353. B	354. D	355. A	356. B	357. B	358. A	359. B	360. C
361. D	362. D	363. A	364. A	365. A	366. C	367. B	368. C	369. C	370. D
371. C	372. D	373. B	374. D	375. D	376. D	377. D	378. C	379. A	380. B
381. D	382. D	383. B	384. C	385. D	386. D	387. B	388. C	389. C	390. D
391. D	392. B	393. B	394. A	395. D	396. A	397. D	398. A	399. A	400. A
401. A	402. C	403. B	404. C	405. A	406. A	407. A	408. A	409. C	410. D
411. B	412. D	413. B							

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